  
 **Student Learning Reflection & Personalised Learning Checklist**

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| **Subject/Course:** | **Separate Chemistry - Foundation** |
| **Student Name:** |  |

**PAPER 1:**

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|  |  | Self Assessment | | |
| Topic | Key knowledge/skills | Red | Amber | Green |
| 4.1.1 – A simple model of the atom, symbols, relative atomic mass, electronic charge and isotopes | State that everything is made of atoms and recall what they are |  |  |  |
| Describe what elements and compounds are |  |  |  |
| State that elements and compounds are represented by symbols; and use chemical symbols and formulae to represent elements and compounds |  |  |  |
| Write word equations and balanced symbol equations for chemical reactions, including using appropriate state symbols |  |  |  |
| Describe what a mixture is |  |  |  |
| Name and describe the physical processes used to separate mixtures and suggest suitable separation techniques |  |  |  |
| Describe how the atomic model has changed over time due to new experimental evidence, inc discovery of the atom and scattering experiments (inc the work of James Chadwick) |  |  |  |
| Describe the difference between the plum pudding model of the atom and the nuclear model of the atom |  |  |  |
| State the relative charge of protons, neutrons and electrons and describe the overall charge of an atom |  |  |  |
| State the relative masses of protons, neutrons and electrons and describe the distribution of mass in an atom |  |  |  |
| Calculate the number of protons, neutrons and electrons in an atom when given its atomic number and mass number |  |  |  |
| Describe isotopes as atoms of the same element with different numbers of neutrons |  |  |  |
| Define the term relative atomic mass and why it takes into account the abundance of isotopes of the element |  |  |  |
| Calculate the relative atomic mass of an element given the percentage abundance of its isotopes |  |  |  |
| Describe how electrons fill energy levels in atoms, and represent the electron structure of elements using diagrams and numbers |  |  |  |
| 4.1.2 – The periodic table | Recall how the elements in the periodic table are arranged |  |  |  |
| Describe how elements with similar properties are placed in the periodic table |  |  |  |
| Explain why elements in the same group have similar properties and how to use the periodic table to predict the reactivity of elements |  |  |  |
| Describe the early attempts to classify elements |  |  |  |
| Explain the creation and attributes of Mendeleev's periodic table |  |  |  |
| Identify metals and non-metals on the periodic table, compare and contrast their properties |  |  |  |
| Explain how the atomic structure of metals and non-metals relates to their position in the periodic table |  |  |  |
| Describe noble gases (group 0) and explain their lack of reactivity |  |  |  |
| Describe the properties of noble gases, including boiling points, predict trends down the group and describe how their properties depend on the outer shell of electrons |  |  |  |
| Describe the reactivity and properties of group 1 alkali metals with reference to their electron arrangement and predict their reactions |  |  |  |
| Describe the properties of group 7 halogens and how their properties relate to their electron arrangement, including trends in molecular mass, melting and boiling points and reactivity |  |  |  |
| Describe the reactions of group 7 halogens with metals and non-metals |  |  |  |
| *Chem ONLY: Describe the properties of transition metals and compare them with group 1 elements, including melting points and densities, strength and hardness, and reactivity (for Cr, Mn Fe, Co, Ni & Cu)* |  |  |  |
| 4.2.1 - Chemical bonds, ionic, covalent and metallic | Describe the three main types of bonds: ionic bonds, covalent bonds and metallic bonds in terms of electrostatic forces and the transfer or sharing of electrons |  |  |  |
| Describe how the ions produced by elements in some groups have the electronic structure of a noble gas and explain how the charge of an ion relates to its group number |  |  |  |
| Describe the structure of ionic compounds, including the electrostatic forces of attraction, and represent ionic compounds using dot and cross diagrams |  |  |  |
| Describe the limitations of using dot and cross, ball and stick, two and three-dimensional diagrams to represent a giant ionic structure |  |  |  |
| Work out the empirical formula of an ionic compound from a given model or diagram that shows the ions in the structure |  |  |  |
| Describe covalent bonds and identify different types of covalently bonded substances, such as small molecules, large molecules and substances with giant covalent structures |  |  |  |
| Represent covalent bonds between small molecules, repeating units of polymers and parts of giant covalent structures using diagrams |  |  |  |
| Draw dot and cross diagrams for the molecules of hydrogen, chlorine, oxygen, nitrogen, hydrogen chloride, water, ammonia and methane |  |  |  |
| Deduce the molecular formula of a substance from a given model or diagram in these forms showing the atoms and bonds in the molecule |  |  |  |
| Describe the arrangement of atoms and electrons in metallic bonds and draw diagrams the bonding in metals |  |  |  |
| 4.2.2 - How bonding and structure are related to the properties of substances | Name the three States of matter, identify them from a simple model and state which changes of state happen at melting and boiling points |  |  |  |
| Explain changes of state using particle theory and describe factors that affect the melting and boiling point of a substance |  |  |  |
| Recall what (s), (l), (g) and (aq) mean when used in chemical equations and be able to use them appropriately |  |  |  |
| Explain how the structure of ionic compounds affects their properties, including melting and boiling points and conduction of electricity (sodium chloride structure only) |  |  |  |
| Explain how the structure of small molecules affects their properties |  |  |  |
| Explain how the structure of polymers affects their properties |  |  |  |
| Explain how the structure of giant covalent structures affects their properties |  |  |  |
| Explain how the structure of metals and alloys affects their properties, including explaining why they are good conductors |  |  |  |
| Explain why alloys are harder than pure metals in terms of the layers of atoms |  |  |  |
| 4.2.3 – Structure and bonding of carbon | Explain the properties of graphite, diamond and graphene in terms of their structure and bonding |  |  |  |
| Describe the structure of fullerenes, and their uses, including Buckminsterfullerene and carbon nanotubes |  |  |  |
| 4.2.4 – Bulk and surface properties of matter including nanoparticles | *Chem ONLY: Compare the dimensions of nanoparticles to other particles and explain the effect of their surface area to volume ratio on their properties* |  |  |  |
| *Chem ONLY: Discuss the applications of nanoparticles and their advantages and disadvantages, including uses in medicine, cosmetics, fabrics and the development of catalysts* |  |  |  |
| 4.3.1 Chemical measurements, conservation of mass and the quantitative interpretation | State that mass is conserved and explain why, including describing balanced equations in terms of conservation of mass |  |  |  |
| Explain the use of the multipliers in equations in normal script before a formula and in subscript within a formula |  |  |  |
| Describe what the relative formula mass (Mr) of a compound is and calculate the relative formula mass of a compound, given its formula |  |  |  |
| Calculate the relative formula masses of reactants and products to prove that mass is conserved in a balanced chemical equation |  |  |  |
| Explain observed changes of mass during chemical reactions in non-enclosed systems using the particle model when given the balanced symbol equation |  |  |  |
| Explain why whenever a measurement is made there is always some uncertainty about the result obtained |  |  |  |
| 4.3.2 – Use of amount of substance in relation to masses of pure substances | Calculate the mass of solute in a given volume of solution of known concentration in terms of mass per given volume of solution |  |  |  |
| 4.3.3 - Yield and atom economy of chemical reactions | *Chem ONLY: Explain why it is not always possible to obtain the calculated or expected amount of a product* |  |  |  |
| *Chem ONLY: Calculate the theoretical amount of a product and percentage yield of a product using the formula % yield = mass of product made/max theoretical mass of product x 100* |  |  |  |
| *Chem ONLY: Describe atom economy as a measure of the amount of reactants that end up as useful products* |  |  |  |
| *Chem ONLY: Calculate the percentage atom economy of a reaction to form a desired product using the equation % atom economy =RfM of desired product/sum of RfM of all reactants x 100* |  |  |  |
| 4.4.1 - Reactivity of metals | Describe how metals react with oxygen and state the compound they form, define oxidation and reduction |  |  |  |
| Describe the arrangement of metals in the reactivity series, including carbon and hydrogen, and use the reactivity series to predict the outcome of displacement reactions |  |  |  |
| Recall and describe the reactions, if any, of potassium, sodium, lithium, calcium, magnesium, zinc, iron and copper with water or dilute acids |  |  |  |
| Relate the reactivity of metals to its tendency to form positive ions and be able to deduce an order of reactivity of metals based on experimental results |  |  |  |
| Recall what native metals are and explain how metals can be extracted from the compounds in which they are found in nature by reduction with carbon |  |  |  |
| Evaluate specific metal extraction processes when given appropriate information and identify which species are oxidised or reduced |  |  |  |
| 4.4.2 – Reactions of Acids | Explain that acids can be neutralised by alkalis, bases and metal carbonates and list the products of each of these reactions |  |  |  |
| Predict the salt produced in a neutralisation reaction based on the acid used and the positive ions in the base, alkali or carbonate and use the formulae of common ions to deduce the formulae of the salt |  |  |  |
| Describe how soluble salts can be made from acids and how pure, dry samples of salts can be obtained |  |  |  |
| ***Required practical 1:*** *preparation of a pure, dry sample of a soluble salt from an insoluble oxide or carbonate using a Bunsen burner to heat dilute acid and a water bath or electric heater to evaporate the solution* |  |  |  |
| Recall what the pH scale measures and describe the scale used to identify acidic, neutral or alkaline solutions |  |  |  |
| Define the terms acid and alkali in terms of production of hydrogen ions or hydroxide ions (in solution), define the term base |  |  |  |
| Describe the use of universal indicator to measure the approximate pH of a solution and use the pH scale to identify acidic or alkaline solutions |  |  |  |
| *Chem ONLY: Describe how to carry out titrations using strong acids and strong alkalis only (sulfuric, hydrochloric and nitric acids to find the reacting volumes accurately* |  |  |  |
| *Chem ONLY:* ***Required practical 2:*** *determination of the reacting volumes of solutions of a strong acid and a strong alkali by titration* |  |  |  |
| 4.4.3 - Electrolysis | Describe how ionic compounds can conduct electricity when dissolved in water and describe these solutions as electrolytes |  |  |  |
| Describe the process of electrolysis |  |  |  |
| Describe the electrolysis of molten ionic compounds and predict the products at each electrode of the electrolysis of binary ionic compounds |  |  |  |
| Explain how metals are extracted from molten compounds using electrolysis and use the reactivity series to explain why some metals are extracted with electrolysis instead of carbon |  |  |  |
| Describe the electrolysis of aqueous solutions and predict the products of the electrolysis of aqueous solutions containing single ionic compounds |  |  |  |
| ***Required practical 3:*** *investigate what happens when aqueous solutions are electrolysed using inert electrodes* |  |  |  |
| 4.5.1 - Exothermic and endothermic reactions | Describe how energy is transferred to or from the surroundings during a chemical reaction |  |  |  |
| Explain exothermic and endothermic reactions on the basis of the temperature change of the surroundings and give examples of everyday uses |  |  |  |
| ***Required practical 4****: investigate the variables that affect temperature changes in reacting solutions* |  |  |  |
| Describe what the collision theory is and define the term activation energy |  |  |  |
| Interpret and draw reaction profiles of exothermic and endothermic reactions, inc identifying the relative energies of reactants and products, activation energy and overall energy change |  |  |  |
| 4.5.2 - Chemical cells and fuel cells | *Chem ONLY: Describe what a simple cell and a battery is and how they produce electricity* |  |  |  |
| *Chem ONLY: Describe why alkaline batteries are non-rechargeable, state why some cells are rechargeable and evaluate the use of cells* |  |  |  |
| *Chem ONLY: Describe fuel cells and compare fuel cells to rechargeable cells and batteries* |  |  |  |
| *Chem ONLY: Describe the overall reaction in a hydrogen fuel cell* |  |  |  |

**PAPER 2:**

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|  |  | Self Assessment | | |
| Topic | Key knowledge/skills | Red | Amber | Green |
| 4.6.1 - Rate of reaction | Calculate the rate of a chemical reaction over time, using either the quantity of reactant used or the quantity of product formed, measured in g/s or cm3/s |  |  |  |
| Draw and interpret graphs showing the quantity of product formed or reactant used up against time and use the tangent to the graph as a measure of the rate of reaction |  |  |  |
| Describe how different factors affect the rate of a chemical reaction, including the concentration, pressure, surface area, temperature and presence of catalysts |  |  |  |
| ***Required practical 5:*** *investigate how changes in concentration affect the rates of reactions by a method involving measuring the volume of a gas produced, change in colour or turbidity* |  |  |  |
| Use collision theory to explain changes in the rate of reaction, including discussing activation energy |  |  |  |
| Describe the role of a catalyst in a chemical reaction and state that enzymes are catalysts in biological systems |  |  |  |
| Draw and interpret reaction profiles for catalysed reactions |  |  |  |
| 4.6.2 - Reversible reactions and dynamic equilibrium | Explain what a reversible reaction is, including how the direction can be changed and represent it using symbols: A + B ⇌ C + D |  |  |  |
| Explain that, for reversible reactions, if a reaction is endothermic in one direction, it is exothermic in the other direction |  |  |  |
| Describe the State of dynamic equilibrium of a reaction as the point when the forward and reverse reactions occur at exactly the same rate |  |  |  |
| 4.7.1 - Carbon compounds as fuels and feedstock | Describe what crude oil is and where it comes from, including the basic composition of crude oil and the general chemical formula for the alkanes |  |  |  |
| State the names of the first four members of the alkanes and recognise substances as alkanes from their formulae |  |  |  |
| Describe the process of fractional distillation, state the names and uses of fuels that are produced from crude oil by fractional distillation |  |  |  |
| Describe trends in the properties of hydrocarbons, including boiling point, viscosity and flammability and explain how their properties influence how they are used as fuels |  |  |  |
| Describe and write balanced chemical equations for the complete combustion of hydrocarbon fuels |  |  |  |
| Describe the process of cracking and state that the products of cracking include alkanes and alkenes and describe the test for alkenes |  |  |  |
| Balance chemical equations as examples of cracking when given the formulae of the reactants and products |  |  |  |
| Explain why cracking is useful and why modern life depends on the uses of hydrocarbons |  |  |  |
| 4.7.2 - Reactions of alkenes and alcohols | *Chem ONLY: State the names and draw structural formulae of the first four members of the alkenes and recognise substances as alkenes from their formulae* |  |  |  |
| *Chem ONLY: Describe the basic composition of alkenes, including the C=C functional group, the general chemical formula for the alkanes and describe what unsaturated means* |  |  |  |
| *Chem ONLY: Describe the combustion reactions of alkenes and the reactions of alkenes with hydrogen, water and the halogens* |  |  |  |
| *Chem ONLY: Draw fully displayed structural formulae of the first four members of the alkenes and the products of their addition reactions with hydrogen, water, chlorine, bromine and iodine* |  |  |  |
| *Chem ONLY: State the functional group of alcohols and the first four members of the homologous series of alcohols and represent alcohols using formulae* |  |  |  |
| *Chem ONLY: Describe some properties and reactions of the first four members of alcohols, including dissolving in water, reacting with sodium, burning in air, oxidation and uses* |  |  |  |
| *Chem ONLY: State the functional group of carboxylic acids and the first four members of the homologous series of carboxylic acids and represent them using diagrams and formulae* |  |  |  |
| *Chem ONLY: Describe some properties and reactions of carboxylic acids, including dissolving in water, reacting with carbonates and reacting with alcohols* |  |  |  |
| 4.7.3 - Synthetic and naturally occurring polymers | *Chem ONLY: Describe how alkenes can be used to make polymers by addition polymerisation* |  |  |  |
| *Chem ONLY: Identify addition polymers and monomers from diagrams and from the presence of the functional group and draw diagrams to represent the formation of an addition polymers* |  |  |  |
| *Chem ONLY: Describe DNA as a large molecule of two polymer chains made from four different monomers called nucleotides in the form of a double helix* |  |  |  |
| *Chem ONLY: State and describe some other naturally occurring polymers such as proteins, starch and cellulose* |  |  |  |
| 4.8.1 - Purity, formulations and chromatography & 4.8.2 - Identification of gases | Define a pure substance and identify pure substances and mixtures from data about melting and boiling points |  |  |  |
| Describe a formulation and identify formulations given appropriate information |  |  |  |
| Describe chromatography, including the terms stationary phase and mobile phase and identify pure substances using paper chromatography |  |  |  |
| Explain what the Rf value of a compound represents, how the Rf value differs in different solvents and interpret and determine Rf values from chromatograms |  |  |  |
| ***Required practical 6:*** *investigate how paper chromatography can be used to separate and tell the difference between coloured substances (inc calculation of Rf values)* |  |  |  |
| Explain how to test for the presence of hydrogen, oxygen, carbon dioxide and chlorine |  |  |  |
| 4.8.3 - Identification of ions by chemical and spectroscopic means | *Chem ONLY: Identify some metal ions from the results of flame tests and describe how to conduct a flame test* |  |  |  |
| *Chem ONLY: Describe how sodium hydroxide solution can be used to identify some metal ions and identify metal ions from the results of their reactions with sodium hydroxide solution* |  |  |  |
| *Chem ONLY: Write balanced equations for the reactions between sodium hydroxide solution and some metal ions to produce insoluble hydroxides* |  |  |  |
| *Chem ONLY: Describe how to identify carbonates using limewater* |  |  |  |
| *Chem ONLY: Describe how to identify negative ions, including halide ions using silver nitrate and sulfate ions using barium chloride* |  |  |  |
| ***Required practical 7:*** *use of chemical tests to identify the ions in unknown single ionic compounds* |  |  |  |
| *Chem ONLY: State the advantages of using instrumental methods to identify elements and compounds compared to chemical tests* |  |  |  |
| *Chem ONLY: Describe the process of and how to use flame emission spectroscopy to identify metal ions; interpret the results of a flame emission spectroscopy tests* |  |  |  |
| 4.9.1 - The composition and evolution of the Earth's atmosphere | Describe the composition of gases in the Earth's atmosphere using percentages, fractions or ratios |  |  |  |
| Describe how early intense volcanic activity may have helped form the early atmosphere and how the oceans formed |  |  |  |
| Explain why the levels of carbon dioxide in the atmosphere changes as the oceans were formed |  |  |  |
| State the approximate time in Earth's history when algae started producing oxygen and describe the effects of a gradually increasing oxygen level |  |  |  |
| Explain the ways that atmospheric carbon dioxide levels decreased |  |  |  |
| 4.9.2 - Carbon dioxide and methane as greenhouse gases | Name some greenhouse gases and describe how they cause an increase in Earth's temperature |  |  |  |
| List some human activities that produce greenhouse gases |  |  |  |
| Evaluate arguments for and against the idea that human activities cause a rise in temperature that results in global climate change |  |  |  |
| State some potential side effects of global climate change, including discussing scale, risk and environmental implications |  |  |  |
| Define the term carbon footprint and list some actions that could reduce the carbon footprint |  |  |  |
| 4.9.3 - Common atmospheric pollutants and their sources | Describe the combustion of fuels as a major source of atmospheric pollutants and name the different gases that are released when a fuel is burned |  |  |  |
| Predict the products of combustion of a fuel given appropriate information about the composition of the fuel and the conditions in which it is used |  |  |  |
| Describe the properties and effects of carbon monoxide, sulfur dioxide and particulates in the atmosphere |  |  |  |
| Describe and explain the problems caused by increased amounts of these pollutants in the air |  |  |  |
| 4.10.1 - Using the Earth's resources and obtaining potable water | State what humans use Earth's resources for, give some examples of natural resources that they use |  |  |  |
| Define the term finite and distinguish between finite and renewable resources |  |  |  |
| Explain what sustainable development is and discuss the role chemistry plays in sustainable development, including improving agricultural and industrial processes |  |  |  |
| State examples of natural products that are supplemented or replaced by agricultural and synthetic products |  |  |  |
| Discuss the importance of water quality for human life, including defining potable water |  |  |  |
| Describe methods to produce potable water, including desalination of salty water or sea water and the potential problems of desalination |  |  |  |
| ***Required practical 8:*** *analysis and purification of water samples from different sources, including pH, dissolved solids and distillation.* |  |  |  |
| Describe waste water as a product of urban lifestyles and industrial processes that includes organic matter, harmful microbes and harmful chemicals |  |  |  |
| Describe the process of sewage treatment and compare the ease of obtaining potable water from waste water as opposed to ground or salt water |  |  |  |
| 4.10.2 - Life cycle assessment and recycling | Describe, carry out and interpret a simple comparative life cycle assessment (LCA) of materials or products |  |  |  |
| Discuss the advantages and disadvantages of LCAs |  |  |  |
| Carry out simple comparative LCAs for shopping bags made from plastic and paper |  |  |  |
| Discuss how to reduce the consumption of raw resources and explain how reusing and recycling reduces energy use (inc environmental impacts) |  |  |  |
| 4.10.3 - Using materials | *Chem ONLY: Define corrosion and describe rusting as an example of corrosion* |  |  |  |
| *Chem ONLY: Describe ways to prevent corrosion, including providing coatings, sacrificial protection and explain how sacrificial protection works* |  |  |  |
| *Chem ONLY: Describe the following alloys bronze, gold, steels and aluminium, their uses and describe the benefits of using alloys instead of pure metals* |  |  |  |
| *Chem ONLY: Compare the properties of materials, including glass and clay ceramics, polymers and composites and explain how their properties are related to their uses* |  |  |  |
| *Chem ONLY: Discuss the different types of polymers and how their composition affects their properties, including thermosoftening and thermosetting polymers* |  |  |  |
| *Chem ONLY: Explain what composites are and provide examples of composites and their benefits over other types of materials* |  |  |  |
| 4.10.4 - The Haber process and the use of NPK fertilisers | *Chem ONLY: Describe the Haber process, including the reactants and products, recycling of remaining hydrogen and nitrogen and the chemical equation* |  |  |  |
| *Chem & HT ONLY: For the Haber process interpret graphs of reaction conditions versus rate* |  |  |  |
| *Chem ONLY: Apply the principles of dynamic equilibrium to the Haber process and discuss the trade-off between the rate of production and the position of equilibrium* |  |  |  |
| *Chem ONLY: Explain how the commercially used conditions for the Haber process are related to the availability and cost of raw materials and energy supplies* |  |  |  |
| *Chem ONLY: Recall the names of the salts produced when phosphate rock is treated with nitric acid, sulfuric acid and phosphoric acid* |  |  |  |
| *Chem ONLY: Describe NPK fertilisers and the compounds they are composed of and compare the industrial production of fertilisers with the laboratory preparations* |  |  |  |